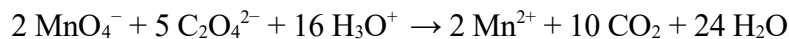
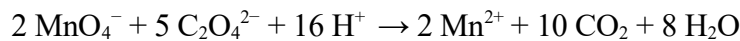
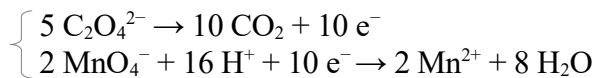
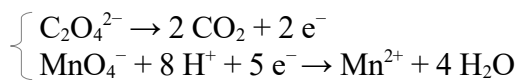
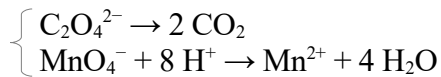
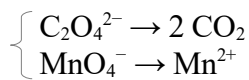
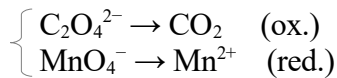
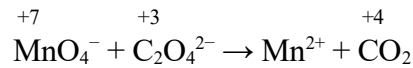
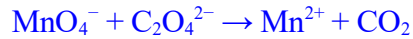


EJEMPLO 6: Calcula la concentración de una disolución de oxalato de potasio, $K_2C_2O_4$, si se necesitan $25,0\text{cm}^3$ de la misma para alcanzar el punto final con $36,7\text{cm}^3$ de una disolución ácida $0,060\text{M}$ de $KMnO_4$. La reacción sin ajustar es:



$$\frac{[C_2O_4^{2-}] \cdot V(C_2O_4^{2-})}{5} = \frac{[MnO_4^-] \cdot V(MnO_4^-)}{2}$$

$$[C_2O_4^{2-}] = \frac{5 \cdot [MnO_4^-] \cdot V(MnO_4^-)}{2 \cdot V(C_2O_4^{2-})}$$

$$[C_2O_4^{2-}] = \frac{5 \cdot 0,060 M \cdot 36,7 mL}{2 \cdot 25,0 mL} = \underline{0,220 M}$$

O también:

$$n(MnO_4^-) = [MnO_4^-] \cdot V(MnO_4^-) = 0,060 M \cdot 0,0367 L = 2,202 \cdot 10^{-3} mol$$

$$2,202 \cdot 10^{-3} mol MnO_4^- \cdot \frac{5 mol C_2O_4^{2-}}{2 mol MnO_4^-} = 5,51 \cdot 10^{-3} mol C_2O_4^{2-}$$

$$[C_2O_4^{2-}] = \frac{n(C_2O_4^{2-})}{V(C_2O_4^{2-})} = \frac{5,51 \cdot 10^{-3} mol}{0,025 L} = \underline{0,220 M}$$