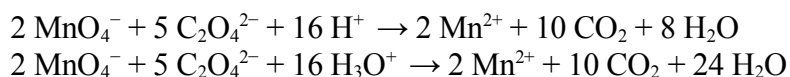
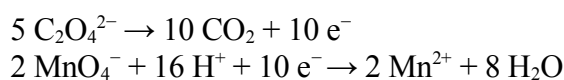
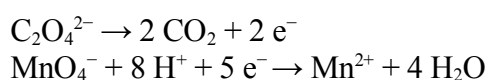
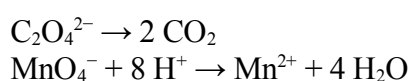
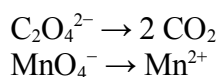
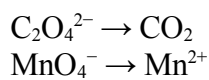
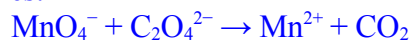


Problema715: Calcula la concentración de una disolución de oxalato de potasio,  $K_2C_2O_4$ , si se necesitan 25,4ml de la misma para alcanzar el punto final con 42,7ml de una disolución ácida 0,080M de  $KMnO_4$ . La reacción sin ajustar es:



$$\frac{[C_2O_4^{2-}] \cdot V(C_2O_4^{2-})}{5} = \frac{[MnO_4^-] \cdot V(MnO_4^-)}{2}$$

$$[C_2O_4^{2-}] = \frac{5 \cdot [MnO_4^-] \cdot V(MnO_4^-)}{2 \cdot V(C_2O_4^{2-})}$$

$$[C_2O_4^{2-}] = \frac{5 \cdot 0,080 M \cdot 0,0427 L}{2 \cdot 0,0254 L} = \underline{\underline{0,336M}}$$