

Problema 447: Calcula la composición centesimal del Na_2CO_3

$$M_m(\text{Na}_2\text{CO}_3) = 2 \cdot 23\text{g} + 12\text{g} + 3 \cdot 16\text{g} = 106\text{g}$$

$$\% \text{Na} = \frac{m_{\text{Na}}}{M_m} \cdot 100 = \frac{2 \cdot 23\text{g}}{106\text{g}} \cdot 100 = \underline{\underline{43,4\% \text{Na}}}$$

$$\% \text{C} = \frac{m_{\text{C}}}{M_m} \cdot 100 = \frac{12\text{g}}{106\text{g}} \cdot 100 = \underline{\underline{11,3\% \text{C}}}$$

$$\% \text{O} = \frac{m_{\text{O}}}{M_m} \cdot 100 = \frac{3 \cdot 16\text{g}}{106\text{g}} \cdot 100 = \underline{\underline{45,3\% \text{O}}}$$

Comprueba que la suma de los porcentajes debe dar 100%

$$\text{Suma} = 43,4 + 11,3 + 45,3 = 100\%$$